

Chemistry

Why Detect Mercury?

Environment:

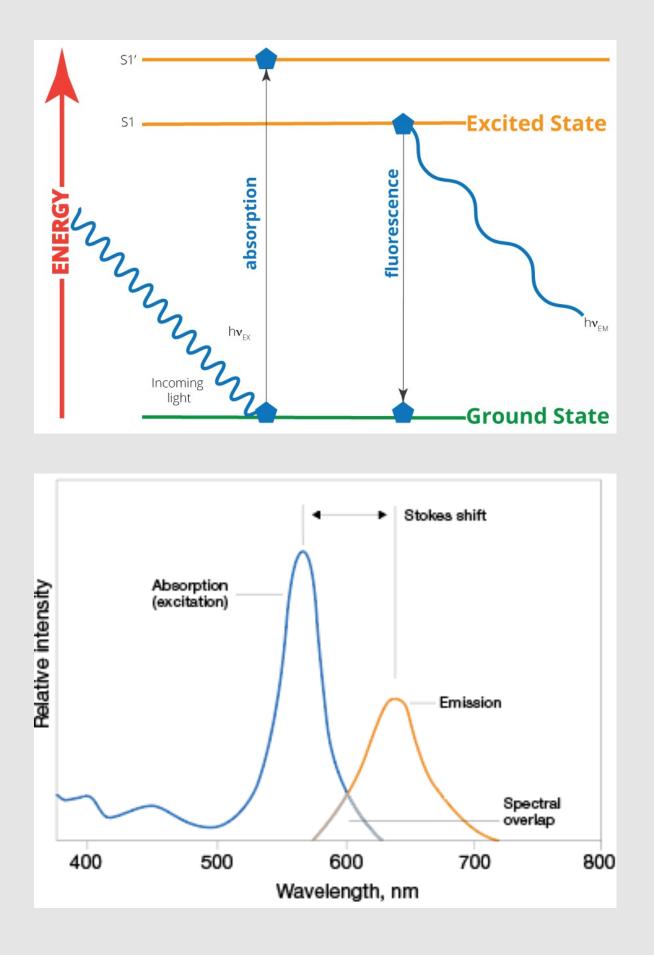
- Mercury is a type of toxic metal that comes in various forms in the environment.
- Amount of mercury in the environment has skyrocketed since since industrialization
- Estimated ~70,000 tons of mercury in the oceans alone!

Effects:

- High doses cause Minamata Disease
- Exposure can harm the brain, heart, kidneys, lungs, and immune system
- Causes nervous system development issues in utero
- Loss of neurons with reactive proliferation of glial cells
- Leads to vascular congestion, petechial hemmorhage, and edema

How Does Fluorescence Work?

- Fluorescence occurs when a compound is irradiated with some frequency of energy (λ). Once excited, the relaxion of the molecule emits another wavelength. This emission is the molecule's fluorescence



References

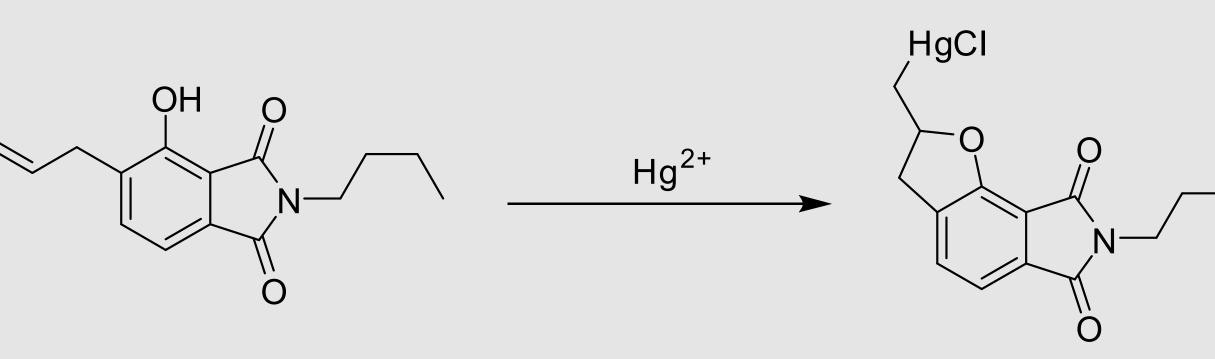
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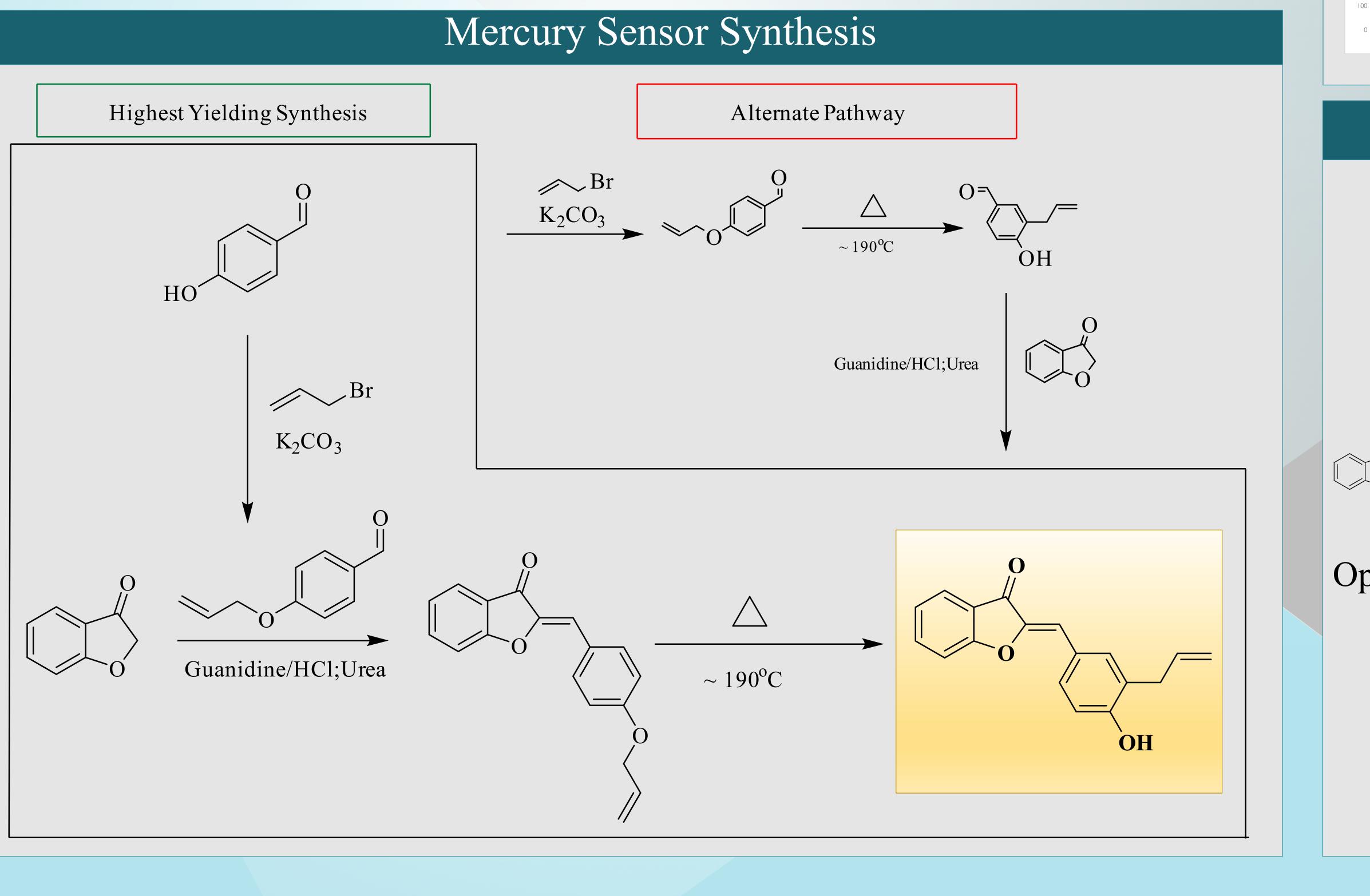
Exploration of Aurone System in Mercury Sensing

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General Oxymercuration Reaction



MS7, A reported Hg²⁺ fluorescent probe designed by Yang, B. et Al.



Reaction Properties:

- Intramolecular
- Speedy reaction time
- Terminal alkene selectivity
- Fluorescent "OFF-Switch"

Possible Issues:

- Interaction with other cations (e.g. Ca^{2+} , Zn^{2+} , Ag^{2+})
- Insignificant drop in fluorescence intensity

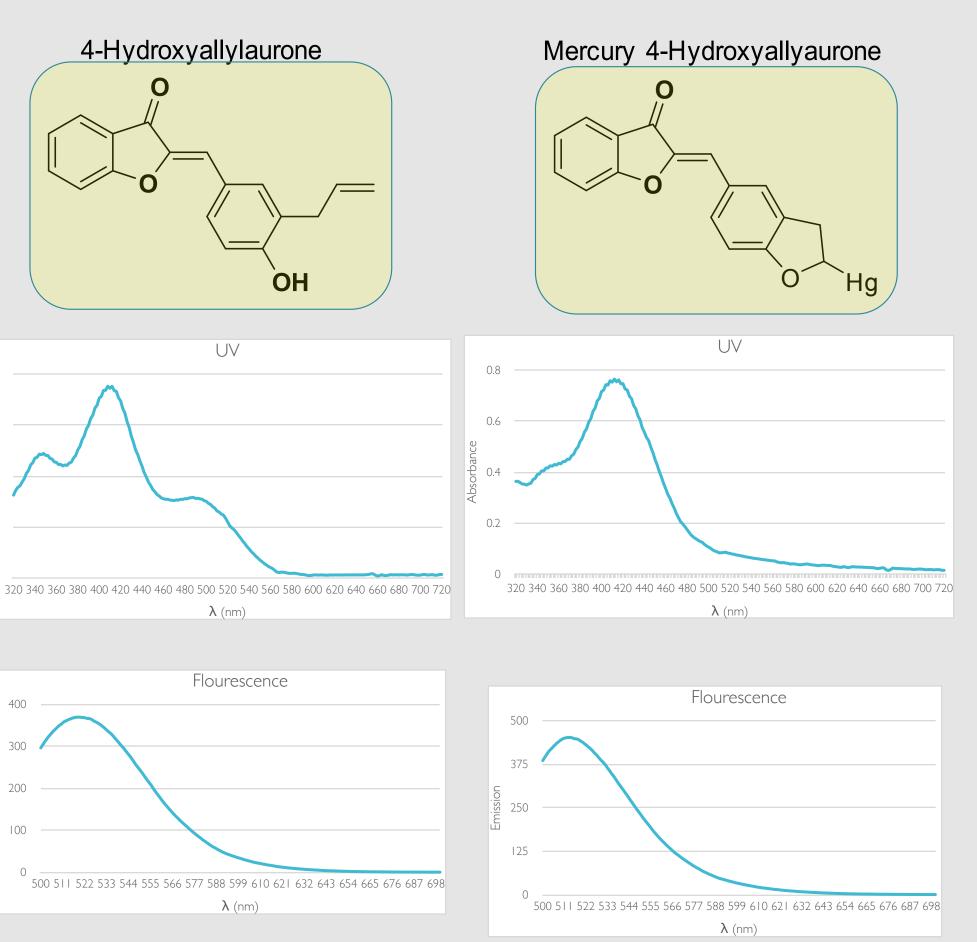
Acknowledgments

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Department of Chemistry



Conclusion



Future Guanidine/HCl:Ur Highest Yield? ~ 190°C K₂CO₃ Optimization: $HgCl_2$ - A shift in fluorescence - Intramolecular H-bond - Hypothesized greater drecrease in flourescence intensity - Potential for solvent solubility manipulation